

- List the different sections of the lab write-up.
- LIST all of the safety issues discussed in class.
- LIST all of the ways to properly use a balance.
- LIST the solutions used and produced during the copper lab; as well as, color, pH, and ie.
- Perform an atom inventory for the following compounds:
 - $C_6H_{12}O_6$
 - $Ca(NO_3)_2$
- What is the correct chemical formula for the following:
 - Calcium chloride
 - Sodium chloride
 - Potassium sulfide
- What is the correct chemical formula for the following:
 - Dinitrogen trioxide
 - disulfur trifluoride
 - Know how to convert the following:
 - kilograms to grams
 - grams to milligrams
- How can you remove a salt from a solution?
- What is the correct diagram (shape) of water?
- Know common elements and their symbols.
- LIST physical and chemical properties of alkali metals.
- What is conserved in a balanced chemical reaction?
- Describe what coefficients are, and why they are used when balancing a chemical reaction?
- What safety precautions are needed when you use the following liquids:
 - NaOH
 - HCl
 - H_2SO_4
- What is the density equation (you will have to solve a problem with it)?
- What are the three major groups of the periodic table?
- LIST the different trends of the periodic table?
- How would you determine if an element is reactive or not?
- Describe an ionic compound; what happens to the electrons?
- What are the different elements in the halogen family?
- Describe the following:
 - compound
 - mixture
 - element
 - Ion
 - Heterogeneous
 - Homogeneous
 - Colloid
 - Suspension
 - Aqueous solution
- LIST the physical and chemical properties of non-metals, metals, and metalloids.
- What is formed from a combustion reaction with hydrocarbons?
- Describe what an isomer is and draw some examples.
- Describe the process of filtration; why would you conduct this process in a lab.



"I SEE YOU GOT A 74 IN YOUR CHEMISTRY FINAL, AND STILL YOU CALL YOURSELF AN EXPERT WITNESS."

26. Draw the following laboratory equipment:

- a. Erlenmeyer flask
- b. Graduated cylinder
- c. Pipette
- d. Watch glass
- e. Beaker
- f. Evaporating Dish
- g. Clay triangle
- h. Crucible

27. List the differences between:

- a. unsaturated
- b. saturated
- c. supersaturated

28. Refer to a solubility curve; how would you interpret the data.

- a. **** FIND THE WORKSHEET THAT WAS COMPLETED****



"Our chemistry lab finals went well.
The firemen just left."

While there will be some overlap between the Common Assessment test and your Course Specific exam. Topics covered will include:

29. Describe a fractionating tower, what settles to the bottom?
30. How can you rank hydrocarbons by boiling points?
31. Describe the intermolecular forces of different petroleum substances.
32. What is organic chemistry?
33. Describe what happens to electrons during a covalent bond.
34. What are the different "roots" for organic compounds:
 - a. 1 carbon = meth-
 - b. And so on...
35. Isomers are what...
36. Vertical and horizontal column of the periodic table are called what?
37. Where are the most relative metals located on the periodic table?
38. Be able to obtain information from a periodic table.
39. Describe protons, electrons, and neutrons.
40. An ion is...
41. Write a chemical reaction and list the reactants and products.
42. LIST some physical properties of different materials.
43. List the parts of a homogenous mixture.
44. What is an atom?

Good luck with your finals!
Study hard, and see me if you
need any help getting ready
for your Chemistry final!

