

No formal report

Lab # 21
monday.

Activated Aluminum Half-Life Measurement

Objective:

In this laboratory you'll be given a thin aluminum foil that has been activated in the OSU TRIGA Reactor. Your objective is to familiarize yourself with the laboratory equipment (with Dr. Farsoni's assistance) and to determine the half-life of the radioactive isotope of aluminum using the activated foil.

Procedure:

1. Set up the GM counter and its supporting electronics so that it is ready for use as a scalar (counter).
2. Confirm that the GM is operating properly and collect any preliminary data that you might need.
3. Set the dual counter/timer to the aluminum-lab counting time (hint: perform this sequence on the counter/time to find the preset program: Mode "Config Module" => scroll to "recall" => scroll to "AL-LAB" => Mode "Oper". At this setting, the counter/time will automatically count for 12 seconds, rest for 48 seconds, and count again in this pattern until stopped.)
4. Aluminum foils (1 cm diameter x 56 micron thickness) will be exposed to a thermal neutron flux of $1 \times 10^9 \text{ n s}^{-1} \text{ cm}^{-2}$ for 60 seconds. After a 5-minute cool-down period, the foils will be brought to the lab.
5. When the aluminum source is in place, start the timer and record each 12-second count for a total of at least 15 trials. Note the time (relative to the end of the irradiation time) that your first count begins.
6. Determine the decay constant and calculate the half-life.

Questions:

1. Considering dead-time correction, determine the decay constant (λ) from your graph and calculate the half-life of the appropriate aluminum isotope.
2. Why is it important to include the effects of dead time? - definition - λ dead time of lower units
3. What is the known radio-aluminum half-life and how different (%) is the known from your calculated half-life? 2.25 minutes or 134.48 secs
4. Estimate the activity of the foil (assume pure aluminum) when it emerged from the reactor and at the beginning of your first count. Estimate both the intrinsic and the absolute efficiency of your GM detector.

Equation

$$A_t = \phi N \sigma (1 - e^{-\lambda t}) (e^{-\lambda t_c})$$

Program on the computer

$$\epsilon \rightarrow \left(\frac{r}{S \Omega \epsilon f} \right)$$

A Brief Introduction to Neutron Activation Analysis

Why NAA?

- Archaeology
- Geology
- Environmental Analysis
- Semiconductor
- Epidemiology
- Forensics



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Why NAA?

Applications

- Archaeology
- Geology
- Environmental Analysis
- Semiconductor
- Epidemiology
- Forensics

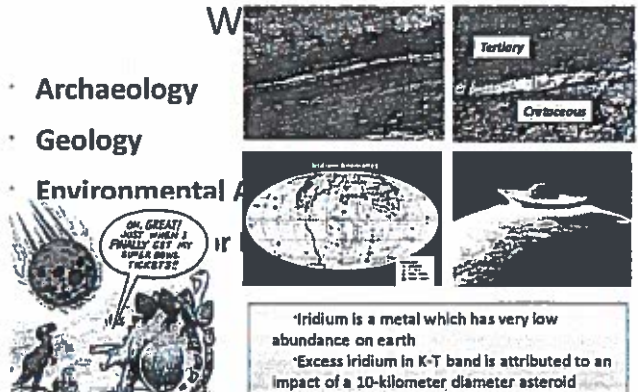
Advantages

- Non-destructive
- Time efficient
- Multi-element
- Not limited by chemical form

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Why NAA?

- Archaeology
- Geology
- Environmental Analysis



Oh, Great! Just when I finally get my Super Bowl tickets!

Iridium is a metal which has very low abundance on earth
Excess Iridium in K-T band is attributed to an impact of a 10-kilometer diameter asteroid

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Why NAA?

Applications

- Archaeology
- Geology
- Environmental Analysis
- Semiconductor
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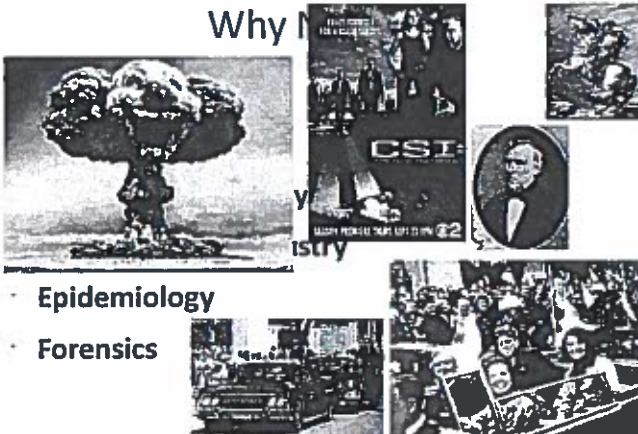
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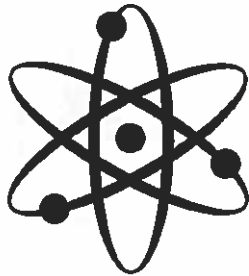
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Why NAA?

- Epidemiology
- Forensics

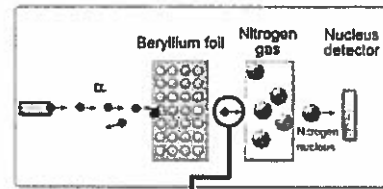


Structure of Atoms



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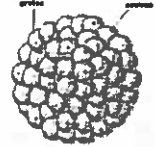
Structure of Atoms



James Chadwick

Carries no charge - neutral
Similar mass as proton

Neutron

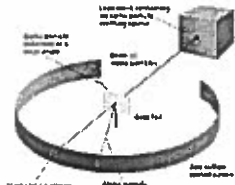


1932

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Structure of Atoms

"It was almost as incredible as if you fired a 15-inch shell at a piece of tissue paper and it came back and hit you."



1911



Ernest Rutherford
1871-1937

Proton
Electron

1919

Hydrogen nucleus is present in other nuclei

Proton

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Nuclear Engineering is the branch of engineering concerned with the application of the breakdown of atomic nuclei and/or other sub-atomic physics, based on the principles of **nuclear physics**.

Chemistry/Chemical Engineering – Atomic level

Nuclear Physics/Nuclear Engineering – Subatomic



Structure of Atoms

"It was almost as incredible as if you fired a 15-inch shell at a piece of tissue paper and it came back and hit you."



1911



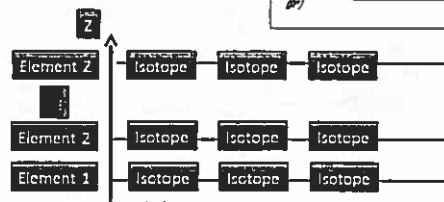
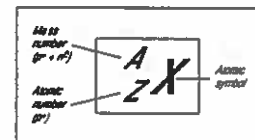
Ernest Rutherford
1871-1937

Proton
Electron

Proton

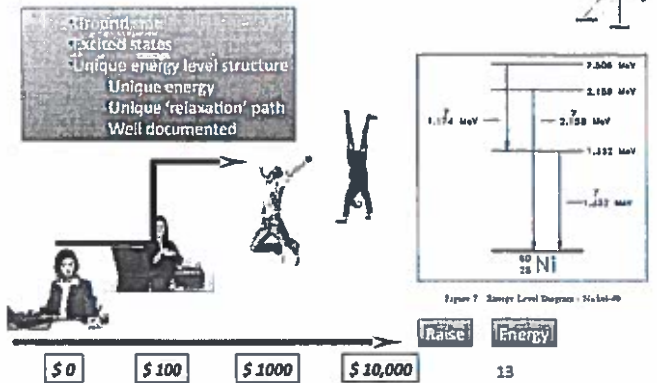
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Isotope vs. Element

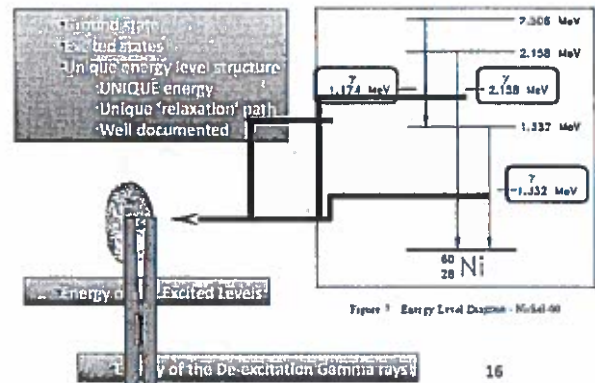


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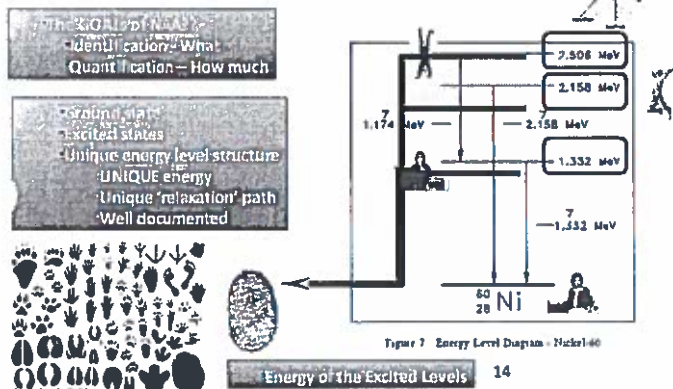
Nucleus Energy Levels



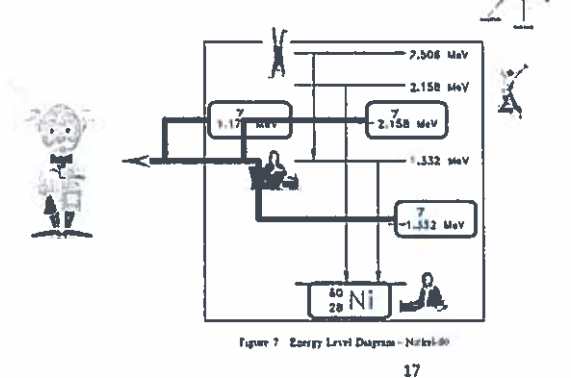
Origin of Gamma-rays



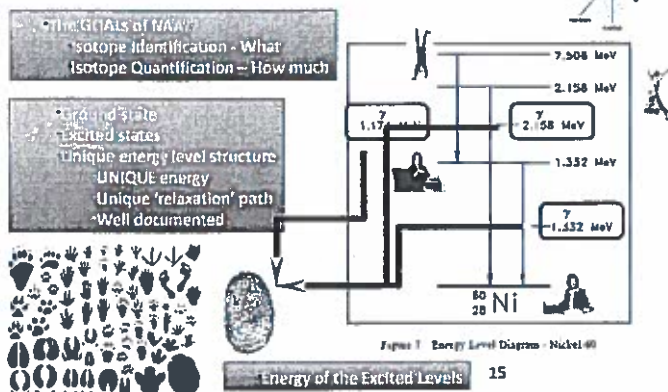
Nucleus Energy Levels



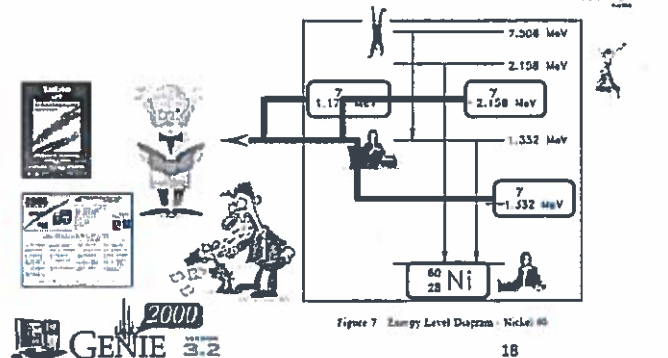
Isotope Identification



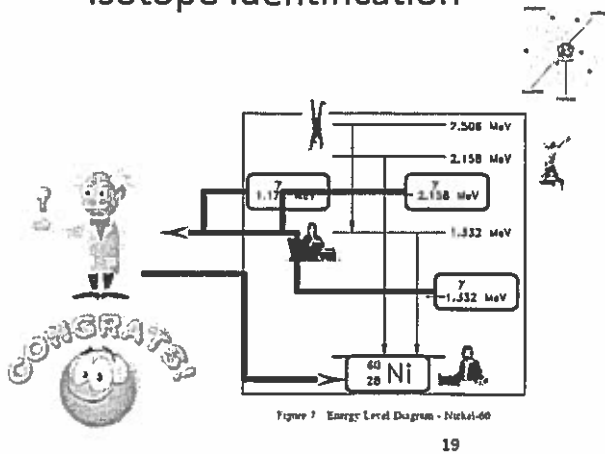
Nucleus Energy Levels



Isotope Identification



Isotope Identification



Al Lab

What is the activity of Al-28 after and irradiation period of t ?

Balance equation of # of Al-28 nuclei

Increase rate:

$$\Phi N \sigma$$

Decrease:

$$\lambda N_d$$

Rate of change:

$$dN_d/dt = \Phi N \sigma - \lambda N_d$$

Initial condition:

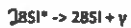
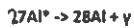
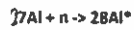
$$N_d(0) = 0$$

Activity

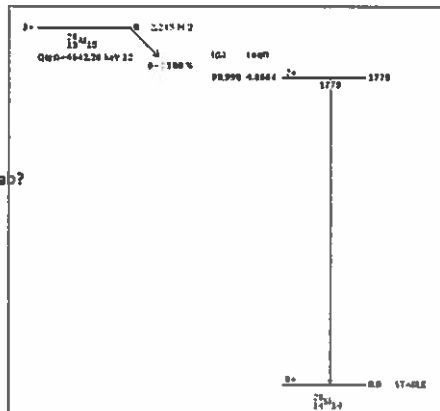
$$A_t = \lambda N_d(t) = \Phi N \sigma (1 - e^{-\lambda t})$$

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Al Lab



What could be measured in the lab?

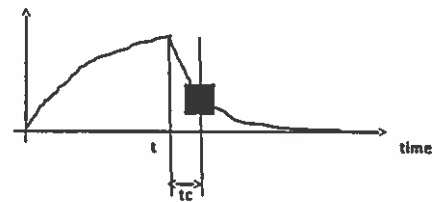


Al Lab

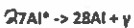
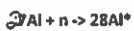
What is the activity of Al-28 after and irradiation period of t ?

$$A_t = \lambda N_d(t) = \Phi N \sigma (1 - e^{-\lambda t})$$

What is the activity after cooling time t_c ?



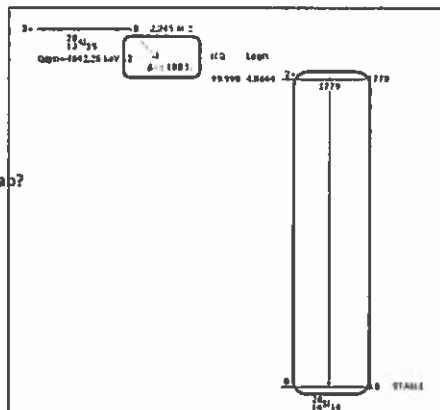
Al Lab



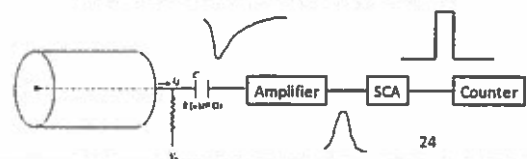
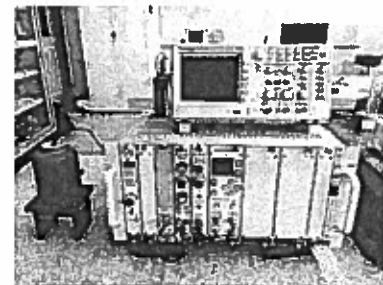
What could be measured in the lab?

- Beta particle
- Gamma ray

What is the half life of Al-28?



Al Lab



$$1 + \frac{\ln z}{1} = \frac{\ln z}{1}$$

1st lab
dont need

Aluminum Lab Worksheet

Background Counts

Background CPM

0

Time	Gross Counts	Δt (min)	Gross CPM	NET CPM
1	1981	0.2	9905	9905
2	1527	0.2	7635	7635
3	1149	0.2	5745	5745
4	923	0.2	4615	4615
5	607	0.2	3035	3035
6	503	0.2	2515	2515
7	399	0.2	1995	1995
8	304	0.2	1520	1520
9	222	0.2	1110	1110
10	166	0.2	830	830
11		0.2	0	0
12		0.2	0	0
13		0.2	0	0
14		0.2	0	0
15		0.2	0	0

^{28}Al Decay Graph

