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### Abstract: Acid Base Titration

The first thing that we did in this experiment was to gather all supplies, chemicals and develop an understanding for the substances being used throughout the experiment. In this particular experiment we were using a titration set-up, along with 0.1 M NaOH, phenolphthalein indicator, a buret, and a KPH solution. The experiment was broken up into two parts: the standardization of NaOH (aq) and then the titration of an unknown acid. We had to first complete the standardization of NaOH in order to find the average molarity so that this number could be used in the calculations of the unknown.

In the standardization of the NaOH, we first measured out 3 samples of KPH that each weighed approximately 0.6 g and documented their individual weights before transferring the substances into three individual flasks. Once all of the samples were in the flasks we added 50 mL of distilled water to dissolve the KPH. Next, we added three drops of phenolphthalein, the indicator substance, to the solution so that the mixture would change color when we titrate the NaOH. Once each of the flasks were ready for titration, we filled the buret with 50mL of NaOH solution and placed the flask under the closed buret system. Finally, we opened the buret, and while the NaOH was entering the flask with the solution we made sure to stir the solution in the flask. Once a pink color started to appear we turned off the buret and kept stirring the mixture. If