

1. Were your values of  $PV$  nearly equal (with a reasonably small percent difference from the average)? Did this aspect of your experimental results seem to verify Boyle's law?
2. Did the data points on your graph of  $P$  vs.  $(1/V)$  fall nearly on a straight line? What does this graphical result imply about the relationship between  $P$  and  $V$ ? Did your graph therefore verify Boyle's law?
3. Would an error be introduced into the experiment if the room temperature should change?
4. Does Boyle's law hold accurately at very-high or very-low temperatures or pressures? Explain.